

Пример письменного коллоквиума

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| 1 | Написать механизм реакции галогенирования на примере хлорирования метана | Инициирование цепи Рост цепи Обрыв цепи |
| 2 | $\text{CH}_3\text{CH}_2\text{CH}_3 + \text{SO}_2 \longrightarrow$ | |
| 3 | $\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3 - \text{C} - \text{H} \\ \\ \text{CH}_3 \end{array} \xrightarrow{\text{HNO}_3}$ | |
| 4 | $\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3 - \text{CH} - \text{CH}_2\text{Br} \end{array} \xrightarrow{\text{Na}}$ | |
| 5 | $\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3 - \text{C} - \text{CH}_3 \\ \\ \text{Br} \end{array} \xrightarrow{\text{Mg/эфир}}$ | |
| 6 | $\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3 - \text{C} - \text{CH}_2 - \text{CH}_3 \\ \\ \text{Cl} \end{array} \xrightarrow{\text{KOH/C}_2\text{H}_5\text{OH}}$ | |
| 7 | $\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3 - \text{C} = \text{CH} - \text{CH}_3 \end{array} \xrightarrow{\text{HBr/H}_2\text{O}_2}$ | |
| 8 | $\begin{array}{c} \text{CH}_3 \quad \quad \text{CH}_3 \\ \quad \quad \\ \text{CH}_3 - \text{CH} - \text{CH} = \text{C} - \text{CH}_3 \end{array} \xrightarrow{\text{KMnO}_4 (\text{p-p})}$ | |
| 9 | $\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3 - \text{C} = \text{C} - \text{CH}_3 \\ \\ \text{CH}_3 \end{array} \xrightarrow{\text{H}_2\text{SO}_4 (\text{ конц.})}$ | |
| 10 | $\text{CH}_2=\text{C}(\text{CH}_3)_2 \xrightarrow{\text{O}_2/\text{Ag}}$ | |
| 11 | $\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3 - \text{C} = \text{CH}_2 \end{array} \xrightarrow{\text{HBr}}$ | |
| 12 | $\text{CH}_3\text{CH}=\text{CH}_2 \xrightarrow{\text{O}_3}$ | |
| 13 | $\text{CH}_2 = \text{CH} - \text{CH} = \text{CH} - \text{CH}_3 \xrightarrow{\text{HBr}}$ | |
| 14 | $\text{CH}_3\text{CH}_2\text{C}\equiv\text{CH} \xrightarrow{\text{HCl}} \xrightarrow{\text{HBr}}$ | |
| 15 | $\text{CH}_3\text{CH}_2\text{C}\equiv\text{CH} \xrightarrow{[\text{Ag}(\text{NO}_3)_2]\text{OH}}$ | |
| 16 | $\text{HC}\equiv\text{CH} + \text{CH}_3-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_2-\text{CH}_3 \longrightarrow$ | |
| 17 | $2 \text{HC}\equiv\text{CH} \xrightarrow{\text{CuCl}/\text{NH}_4\text{Cl}}$ | |